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Catalyst and Process of Paraffin Hydrocarbon Conversion

BACKGROUND OF THE INVENTION AND PRIOR ART

The present invention relates to a process for the conversion of paraffin hydrocarbons catalysed by a mixture of an acidic ionic liquid catalyst and a Brønsted acid (proton donating acid).

Paraffin hydrocarbons with high degree of branching are known to be useful blending components for motor gasoline due to their high octane numbers. Such paraffin hydrocarbon fraction can be produced in an isomerisation process increasing the octane number of the C₄-C₉ cuts. Isomerisation of C₄, C₅ and C₆ paraffins are common refinery processes based on use of e.g. an acidic Friedel-Crafts catalyst such as AlCl₃. Processes including higher fractions (C₇ to C₉ hydrocarbons) meet with significant difficulties due to low selectivity and low octane number of the once-through products.

A relatively new class of acidic catalysts based on ionic liquids, e.g. produced from AlCl₃, has recently been described in the literature (P. Wasserscheid, W. Keim, Angew. Chem., Int. Ed., 2000, V. 39, pages 3772-3789; T. Welton, Chem. Rev., 1999, V. 99, pages 2071-2083). This group of compounds also referred to as molten salts are constituted of:

- (1) an inorganic anion typically formed from metal halides, such as AlCl₄⁻, Al₂Cl₇⁻ or other inorganic anions (SO₄²⁻, NO₃⁻, PF₆⁻, CF₃SO₃⁻, BF₄⁻ etc.), and
- (2) an organic cation typically derived from N-heterocyclic or alkylammonium entities.

The melting point of ionic liquids is relatively low and an increasing number of ionic liquids are described with melting points below room temperature. Below some characteristics of ionic liquids are listed:

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- (1) They have a liquid range of about 300°C.
- (2) They are good solvents for a wide range of inorganic, organic and polymeric materials.
- (3) They exhibit Brønsted and Lewis acidity as well as
10 superacidity.
- (4) They have low or no vapour pressure.
- (5) Most ionic liquids are thermally stable up to near 200°C, some ionic liquids are stable at much higher temperature (about 400-450°C).
- 15 (6) They are relatively cheap and easy to prepare and upscale.
- (7) They are non-flammable and easy in operation.
- (8) They are highly polar but non-coordinating materials.

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Thus, the term "ionic liquid" in the following description shall refer to salts consisting of ions, which exist in the melted form and consist of organic nitrogen-containing heterocyclic or aliphatic cations and inorganic anions.

25

Ionic liquids most frequently demonstrate Lewis acidic properties once they are formed by metal halides. In many cases, however, the ionic liquids show also strong Brønsted (proton) acidity. The proton acidity may originate both
30 from the cation if it contains a proton at the quarternized N atom or from the anion if it contains protons for instance in HSO_4^- , H_2PO_4^- .

Also HCl produced via partial hydrolysis for example of the chloroaluminate anion can explain strong proton acidity of the ionic liquids. Addition of a Brønsted Acid, e.g. H_2SO_4 , to an ionic liquid containing chloroaluminate anions, will
 5 also increase the amount of protons in the medium and in case the Brønsted Acid react with the ionic liquid HCl is liberated to the medium.

Lewis-acidic properties of ionic liquids are governed by
 10 two major factors: (1) the nature of the anion, and (2) the molar ratio of the organic part to the inorganic part (for instance in the case of ionic liquids based on metal halides $\text{Me}(\text{Hal})_n$ by the molar fraction of $\text{Me}(\text{Hal})_n$). If $X_{\text{Me}(\text{Hal})_n} < 0.5$ the ionic liquid is called basic; if $X_{\text{Me}(\text{Hal})_n} = 0.5$
 15 this is the case of neutral ionic liquid, and finally if $X_{\text{Me}(\text{Hal})_n} > 0.5$ the ionic liquid can be classified as acidic or in some cases superacidic.

The effect of superacidity of ionic liquids is quite frequently observed for AlCl_3 -based compositions. Sometimes
 20 this effect is related to the presence of dry HCl in the system, which is dissolved in the ionic liquid. The Hammett function H_0 for such systems ($H_0 = -18$) indicates superacidic properties of the ionic liquids comparable with those of
 25 HF-TaF_5 ($H_0 = -16$) and "magic acid" HF-SbF_5 or $\text{FSO}_3\text{H-SbF}_5$ ($H_0 = -25$). All these systems are much stronger acids as compared to the conventional 100% H_2SO_4 ($H_0 = -12$), which marks the border of superacidity. Such ionic liquids are also
 30 stronger than the solid superacids like SO_4/ZrO_2 ($H_0 = -16$), $\text{H}_3\text{PW}_{12}\text{O}_{40}$ ($H_0 = -13.5$) or H-Nafion ($H_0 = -12$).

SUMMARY OF THE INVENTION

The object of the present invention is to provide an improved catalyst and a process for the conversion of linear and/or branched paraffin hydrocarbons.

Based on the observation that ionic liquid catalyst combined with a Brønsted Acid provides a catalytic composition with improved activity compared to ionic liquid this invention is a catalyst composition for use in a hydrocarbon conversion process with the provision that the hydrocarbon conversion process is not cracking of polymers, which composition comprises

- (a) an ionic liquid catalyst comprised of a N-containing heterocyclic and/or aliphatic organic cation and an inorganic anion derived from metal halides or mixed metal halides, and
- (b) one or more Brønsted Acids.

It has been found that the above catalyst composition is particularly useful in isomerisation of paraffin hydrocarbons.

Consequently, a further aspect of the invention is a process for isomerisation of hydrocarbon feed comprising paraffinic hydrocarbons in the presence of a composite catalyst comprising

- (a) an ionic liquid catalyst comprised of a N-containing heterocyclic and/or aliphatic organic cation and an inorganic anion derived from metal halides or mixed metal halides, and
- (b) one or more Brønsted Acids.

DETAILED DESCRIPTION OF THE INVENTION

The ionic liquids used for preparation of the catalyst composition and the hydrocarbon isomerisation reaction represent salts formed by an organic cation such as N-containing heterocyclic or N-containing aliphatic moiety and an inorganic anion, which may be an anion derived from metal halides or mixed metal halides. The cation may be an alkyl substituted pyridinium, piperidinium, quinolinium (or similar amine compounds) with one or several alkyl or aryl groups or an alkyl ammonium (mono-alkyl, di-alkyl, tri-alkyl or tetra-alkyl ammonium compound). The anion may be derived from any metal halide with strong Lewis acidic properties for instance AlCl_4^- , AlBr_4^- , GaCl_4^- , Al_2Cl_7^- , $\text{Al}_2\text{Cl}_6\text{Br}^-$ and the like. The ionic liquid chosen for paraffin isomerisation may be characterised by the amine: Lewis acid molar ratio from 1:3 to 2:1, more preferably from 1:2.5 to 1:1.

The Brønsted Acid used in combination with the ionic liquids as catalysts can be chosen from HCl , HBr , $\text{CH}_3\text{SO}_3\text{H}$ (and other alkane sulphonic acids), $\text{CH}_3\text{CO}_2\text{H}$ (and other carboxylic acids), $\text{CF}_3\text{SO}_3\text{H}$ (and other fluorinated alkane sulphonic acids), $\text{CF}_3\text{CO}_2\text{H}$ (and other fluorinated carboxylic acids), ClSO_3H , FSO_3H , H_2SO_4 , H_3PO_4 and the like. Physical mixtures of several of these compounds may also be used.

The Brønsted Acid can be added in gaseous, liquid or solid form to the ionic liquid in some cases resulting in the formation of a heterogeneous mixture. Some of the Brønsted Acids react with the ionic liquid liberating HCl (if the ionic liquid is based on e.g. a chloroaluminate compound).

The mixture of ionic liquid and Brønsted Acid can be used as catalyst as such, or it can be treated by appropriate means, e.g. heat treatment.

5 The catalyst composition according to the invention gives a novel strongly acidic catalyst, which is significantly more active than common ionic liquids. As such it can be used in a large number of hydrocarbon conversions, where also room-temperature ionic liquids are used. Among these processes
10 of potential commercial interest are various alkylation, oligomerisation and isomerisation reactions. The list of such possible applications is given in D. Zhao, M. Wu, Y. Kou, E. Min, Catalysis Today, V. 74, 2002, pages 157-189, whose content hereby is incorporated into this patent dis-
15 closure by reference thereto.

The solubility of hydrocarbons in ionic liquids is limited and for instance paraffins and naphthenes are generally immiscible with ionic liquids. Olefins and aromatic compounds
20 demonstrate a clear dependence of the solubility on the oleophilic properties of the ionic liquid. The longer the chain length of the radical attached to the N-heterocyclic moiety, the higher the solubility of olefins and aromatics in the ionic liquids. However, most of the commonly used
25 organic solvents and reagents are immiscible with ionic liquids. This simplifies the use of ionic liquids in a biphasic system and provides a procedure for a simple product/catalyst separation.

30 Paraffin isomerisation can be carried out in pressurised equipment under high pressure or in a glass vessel at atmospheric pressure. The pressure in the autoclave can be

varied from 1 bar to 60 bar. Any gas like helium, argon, nitrogen, hydrogen or dry air can be used in the reaction. The reaction temperature can vary in a range from -30°C to 150°C. Temperatures out of this range can also be used although they are less preferred.

Linear n-paraffins such as n-butane, n-pentane, n-hexane, n-heptane, n-octane, n-nonane and monomethylalkanes such as 2-methylhexane and 3-methylhexane or a mixture thereof can be used as substrates of the isomerisation process forming a product containing paraffin hydrocarbons with a higher degree of branching.

The hydrocarbon feeds used for the isomerisation experiments in this disclosure is specified below.

Experimental Procedures 1-3:

17.7 wt% n-heptane, 21.0 wt% 2-methylhexane, 20.9 wt% 3-methylhexane, 36.7 wt% methylcyclohexane, 1.1 wt% 2,4-dimethylpentane, 1,6 wt% 2,3 dimethylpentane and 1.0 wt% of other C7 isomer compounds.

Experimental Procedure 4:

19.5 wt% n-heptane, 20.4 wt% 2-methylhexane, 20 wt% 3-methylhexane, 35.6 wt% methylcyclohexane, 1 wt% 2,4-dimethylpentane, 1,5 wt% 2,3 dimethylpentane and 2.0 wt% of other C7 isomer compounds.

EXAMPLES

Example 1

In an inert atmosphere (N_2), trimethylamine hydrochloride
5 (39.13 g, 0.409 mole) is added to aluminium chloride (98.28
g, 0.737 mole). The light-brown viscous melt, which forms
are heated to $90^\circ C$ under stirring and kept at this tempera-
ture for 2 hours. From the resulting liquid may precipitate
some solid $AlCl_3$ after cooling to room temperature. In the
10 isomerisation experiments described below only the liquid
phase has been used as catalyst. The ionic liquid can be
stored in inert atmosphere (N_2) without decomposition.

Example 2

15 In an inert atmosphere (N_2), a 2-neck Schlenk flask
equipped with a mechanical stirrer is charged with 30 ml
ionic liquid (42 g) prepared according to Example 1 and 30
ml of the organic hydrocarbon feed. A certain amount of
Brønsted Acid (see Table 1) is added to the mixture. The
20 system is vigorously stirred (700 rpm) at constant tempera-
ture. Samples of the hydrocarbon phase are taken at regular
intervals and analyzed by a gas chromatograph.

Example 3

25 In an inert atmosphere (N_2) a 2-neck Schlenk flask is
charged with 30 ml ionic liquid (42 g) prepared according
to Example 1 and a certain amount of Brønsted Acid (see Ta-
ble 1). This mixture is heated to $90^\circ C$ and left under stir-
ring for 1 hour. After cooling to room temperature, 30 ml
30 of the organic hydrocarbon feed is added to the mixture.
The system is vigorously stirred (700 rpm) using mechanical
agitation at constant temperature. Samples of the hydrocar-

bon phase are taken at regular intervals and analyzed by a gas chromatograph.

Example 4

5 In an inert atmosphere (N_2), an autoclave with mechanical stirrer is charged with 40 ml ionic liquid (56 g) prepared according to Example 1 and 40 ml of the organic hydrocarbon feed. A certain amount of Brønsted acid (see Table 1) is added to the mixture. The system is pressurised with 5 bar
10 helium (for sampling) and afterwards vigorously stirred (700 rpm) at constant temperature. Samples of the hydrocarbon phase are taken at regular intervals and analysed by a gas chromatograph.

15 Example 5

In an inert atmosphere (N_2), a 2-neck Schlenk flask equipped with a mechanical stirrer is charged with 30 ml ionic liquid (42 g) prepared according to Example 1. A stream of HCl gas is bubbled through the ionic liquid for
20 30 min, thereby dissolving HCl in the ionic liquid. 30 ml of the organic hydrocarbon feed, which earlier has been saturated with HCl gas, are added to the ionic liquid. The system is vigorously stirred (700 rpm) at constant temperature. Samples of the hydrocarbon phase are taken at regular
25 intervals and analyzed by a gas chromatograph.

Table 1

DEFINITIONS

Multi-branched C₇ products:
 5 Dimethylpentanes and trimethylbutane.

C₆- products:
 Compounds containing six and less than six carbon atoms.

10 C₈+ products:
 Compounds containing eight and more than eight carbon atoms.

Normalised yield of multi-branched C₇ products is defined as:
 15 $100 \times (\text{sum of multi-branched C}_7 \text{ products}) / (\text{sum of C}_7 \text{ compounds excluding cyclic C}_7 \text{ compounds})$.

Selectivity to multi-branched C₇ products is defined as:
 20 $100 \times (\text{sum of multi-branched C}_7 \text{ products}) / (\text{sum of multi-branched C}_7 \text{ products} + \text{C}_6 \text{ products} + \text{C}_{8+} \text{ products})$.

Example	Brønsted acid	Amount of Brønsted acid (g)	Temperature (°C)	Time (min)	Normalised yield of multi-branched isomers (wt%)	Selectivity (wt%)
2 (a)	None (reference example)		25	30	6.6	97.8
				60	7.6	98.2
				90	8.2	98.6
				120	8.6	98.6
				180	9.5	99.1
				240	10.2	99.3
				300	10.7	99.2
2 (b)	H ₂ SO ₄ (96 wt-%)	2.30	25	30	7.2	98.1
				60	11.9	98.7
				90	17.7	98.4
				120	24.5	96.6
				150	28.2	93.9
				180	29.5	91.2
2 (c)	H ₂ SO ₄ (96 wt-%)	5.52	25	5	5.3	80.8
				10	6.6	98.4
				15	8.1	98.7
				30	11.7	99.1
				60	18.3	98.6

2 (d)	H ₂ SO ₄ (96 wt-%)	6.81	25	30 60 90 120 150 180	8.4 15.5 19.8 27.0 28.3 28.6	98.2 98.0 92.7 90.3 90.4 89.2
2 (e)	CF ₃ SO ₃ H	3.48	25	5 10 15 30 60 150 180 240	5.9 7.3 8.8 11.6 14.6 16.9 17.3 18.0	96.8 98.6 98.8 99.0 99.1 99.0 99.1 99.0
2 (f)	CF ₃ SO ₃ H	6.78	25	30 60 90 120 150 180	9.2 14.3 17.2 19.5 20.0 20.4	98.6 98.8 98.3 98.3 98.1 98.2
2 (g)	CF ₃ SO ₃ H	10.18	25	30 60 90 120 150 180	7.0 7.9 8.2 8.5 8.7 8.9	98.5 98.8 99.0 99.0 98.8 97.0
2 (h)	ClSO ₃ H	0.53	0	30 60 90 120 150 180	5.3 6.3 7.4 8.8 10.8 13.7	97.8 98.4 98.8 98.7 98.2 99.3
2 (i)	ClSO ₃ H	1.40	25	30 60 90 120 150 180	26.4 34.5 35.9 36.3 36.4 36.3	90.0 72.0 69.6 68.1 68.2 66.4
2 (j)	ClSO ₃ H	2.72	25	5 10 15 30 60 120 180	9.2 16.3 23.0 33.3 37.8 38.7 38.5	97.9 96.2 92.4 76.8 66.1 64.0 62.5

2 (k)	H_3PO_4	2.27	25	30	7.5	98.0
				60	10.9	98.7
				90	12.5	97.3
				120	13.2	98.6
				150	13.7	98.5
				180	14.0	98.7
2 (l)	H_3PO_4	4.54	25	30	8.1	97.7
				60	11.2	97.4
				90	12.7	97.9
				120	13.4	97.8
				150	13.8	99.0
				180	14.2	99.1
2 (m)	H_3PO_4	2.27	45	30	24.3	88.9
				60	27.1	88.4
				90	28.0	85.6
				120	28.5	82.8
				150	28.9	81.2
				180	29.2	79.1
3 (a)	ClSO_3H	1.55	25	30	26.2	90.3
				60	34.9	71.5
				90	36.1	69.2
				120	36.9	67.8
3 (b)	H_3PO_4	2.27	25	30	14.6	97.9
				60	18.7	97.3
				90	20.5	97.0
				120	21.7	96.4
				150	23.2	95.3
				180	24.0	93.5
4 (a)	H_2SO_4 (96 wt-%)	2.94	25	30	15.9	96.7
				60	23.3	96.1
				86	27.0	93.8
				140	32.8	80.0
				195	38.2	63.4
				236	40.9	56.8
4 (b)	ClSO_3H	3.5	25	8	10.0	98.1
				15	20.6	95.6
				30	28.4	83.5
				45	35.2	69.1
				60	36.5	65.6
				75	37.5	63.5
				90	38.8	61.2
5 (a)	HCl		25	5	5.2	94.5
				10	6.1	95.2
				15	8.0	95.9
				30	10.2	96.6
				60	13.2	96.8
				120	15.3	91.5
				180	16.7	94.7
				240	17.8	96.4